

Question Number	Acceptable Answers	Reject	Mark
4 (f) (i)	$(7.5 \times 0.81) = 6.075 / 6.08 \text{ (g)}$ Ignore sf except 1 sf	6.07 Wrong units	1

Question Number	Acceptable Answers	Reject	Mark
4 (f) (ii)	<p>Look at final answer. 67% scores 3 marks; answers with 3sf rounding to 67 score 2 marks. If this is incorrect follow this scheme:</p> <p>METHOD 1</p> <p>Mol butan-1-ol = $(6.075/74)$ $= 0.0820945$ (1)</p> <p>maximum mass 1-bromobutane = $(0.0820945 \times 137) = 11.246959 \text{ g}$ (1)</p> <p>% yield = $((7.5/11.24659) \times 100)$ $= 66.85)$ $= 67\% \text{ to 2 sf}$ (1)</p> <p>OR METHOD 2</p> <p>$7.5/137 = 0.0547445 \text{ mol (bromobutane)}$ (1)</p> <p>$6.075/74 = 0.0820945 \text{ mol butan-1-ol}$ (1)</p> <p>% yield = $((0.05474455) \times 100 / 0.0820945)$ $= 66.85)$ $= 67\% \text{ to 2 sf}$ (1)</p> <p>Also TE from one step of the calculation to the next and TE on 4f(i) unless yield > 100%.</p> <p>Use of 6.08 gives 0.082161 mol, 11.256216 g bromobutane, final answer 67%</p> <p>11.3g bromobutane gives 66%.</p>	<p>Percentages calculated from volumes with no conversion to mol or mass.</p> <p>$6.075/7.5 \times 100 = 81\%$ scores 0</p> <p>67.0 (This is 3sf)</p>	3

Total for Question 4 = 15 marks

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DP Hallahan



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