

# Part B-1

Answer all questions in this part.

Directions (31–50): For each statement or question, record on your separate answer sheet the number of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Setting/Chemistry.

31 What is the number of electrons in an  $\text{Al}^{3+}$  ion?

- (1) 10 (3) 3  
(2) 13 (4) 16

Al +13p

32 The valence electron of which atom in the ground state has the greatest amount of energy?

- (1) cesium (3) rubidium  
(2) lithium (4) sodium

33 The numbers of protons and neutrons in each of four different atoms are shown in the table below.

Protons and Neutrons in Four Different Atoms

| Atom | Number of Protons | Number of Neutrons |
|------|-------------------|--------------------|
| A    | 8                 | 8                  |
| D    | 9                 | 9                  |
| E    | 9                 | 10                 |
| G    | 10                | 10                 |

Which two atoms represent isotopes of the same element?

- (1) A and D (3) E and D  
(2) A and G (4) E and G

34 Which elements have the most similar chemical properties?

- (1) boron and carbon  
(2) oxygen and sulfur  
(3) aluminum and bromine  
(4) argon and silicon

35 Which element reacts with oxygen to form ionic bonds?

- (1) calcium (3) chlorine  
(2) hydrogen (4) nitrogen

36 The table below gives the atomic mass and the abundance of the two naturally occurring isotopes of chlorine.

Naturally Occurring Isotopes of Chlorine

| Isotopes         | Atomic Mass of the Isotope (u) | Natural Abundance (%) |
|------------------|--------------------------------|-----------------------|
| $^{35}\text{Cl}$ | 34.97                          | 75.76                 |
| $^{37}\text{Cl}$ | 36.97                          | 24.24                 |

Which numerical setup can be used to calculate the atomic mass of the element chlorine?

- (1)  $(34.97 \text{ u})(75.76) + (36.97 \text{ u})(24.24)$   
(2)  $(34.97 \text{ u})(0.2424) + (36.97 \text{ u})(0.7576)$   
(3)  $(34.97 \text{ u})(0.7576) + (36.97 \text{ u})(0.2424)$   
(4)  $(34.97 \text{ u})(24.24) + (36.97 \text{ u})(75.76)$

37 Which general trends in first ionization energy and electronegativity values are demonstrated by Group 15 elements as they are considered in order from top to bottom?

- (1) The first ionization energy decreases and the electronegativity decreases.  
(2) The first ionization energy increases and the electronegativity increases.  
(3) The first ionization energy decreases and the electronegativity increases.  
(4) The first ionization energy increases and the electronegativity decreases.

38 An aluminum sample has a mass of 80.01 g and a density of 2.70 g/cm<sup>3</sup>. According to the data, to what number of significant figures should the calculated volume of the aluminum sample be expressed?

- (1) 1 (3) 3  
(2) 2 (4) 4

$$D = \frac{M}{V}$$

$$V = \frac{M}{D}$$

$$\frac{80.01 \text{ (4)}}{2.70 \text{ (3)}}$$

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**Barron's Educational Series, Andre,  
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