

## BUILD Vocabulary

**ionic bond** chemical bond formed when one or more electrons are transferred from one atom to another

**ion** atom that has a positive or negative charge

**covalent bond** type of bond between atoms in which the electrons are shared

**molecule** smallest unit of most compounds that displays all the properties of that compound

**van der Waals forces** weak attraction that develops between oppositely charged regions of nearby molecules

**Preface** The prefix *co-* and *co-* mean "together." A compound is made up of elements joined together, and a covalent bond connects atoms that share electrons together. **How are covalent bonds different from ionic bonds?**

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## Chemical Bonds

**KEY QUESTION** What are the main types of chemical bonds?

The atoms in compounds are held together by chemical bonds. Bonds are formed by the electrons in the outer shell, which are called valence electrons. The main types of chemical bonds are ionic bonds and covalent bonds.

**Ionic Bonds** **Ionic bonds** are formed when one or more electrons are transferred from one atom to another. This transfer causes the atom losing the electrons to become positively charged. The atom gaining the electrons becomes negatively charged. The positively and negatively charged atoms become **ions**. The oppositely charged ions are attracted to each other and form an ionic bond.

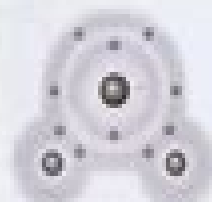
**Covalent Bonds** Sometimes atoms share electrons rather than transferring them. These shared electrons move around the nuclei of both atoms. The bond formed by sharing electrons is called a **covalent bond**. Atoms can share between two and six electrons. The structure that results is known as a **molecule**. A molecule is the smallest unit of most compounds.

**Weak Interactions** The strongest chemical bonds are ionic or covalent. Weak interactions between atoms are also important. Within living organisms, many molecules interact only briefly. **Van der Waals forces** are one type of weak interaction between molecules. They occur when molecules are very close together.

Hydrogen bonds are another type of weak interaction. They occur between a hydrogen atom of one molecule and an oxygen or nitrogen atom of a neighboring molecule.

### Visual Reading Skill Covalent Bonds

1. What kind of chemical bond holds the water molecule together? \_\_\_\_\_
2. What particles are shared between the atoms in the molecule? \_\_\_\_\_
3. What is the other main type of chemical bond? \_\_\_\_\_
4. What is the difference between covalent and ionic bonds? \_\_\_\_\_



Water molecule ( $H_2O$ )

## Properties of Water

**READING TOOL** Cause and Effect As you read the part of the lesson under "The Water Molecule," use the table below to list the causes and effects of the properties of water. Many of the causes may be the same, but each characteristic will have a different effect. Part of the table is completed for you.

Characteristic of Water	Cause	Effect
Hydrogen Bonding		The slightly negative end of a water molecule is attracted to the slightly positive end of another water molecule.
Expands Upon Freezing		
Dissolves Many Substances	Water forms hydrogen bonds.	
Cohesion		
Adhesion		
Heat Capacity		

# Miller Levine Reading And Study Workbook A Answer Key

**AN Whitehead**



## **Miller Levine Reading And Study Workbook A Answer Key:**

## **Miller Levine Reading And Study Workbook A Answer Key** Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the energy of words has be evident than ever. They have the ability to inspire, provoke, and ignite change. Such could be the essence of the book **Miller Levine Reading And Study Workbook A Answer Key**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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### **Table of Contents Miller Levine Reading And Study Workbook A Answer Key**

1. Understanding the eBook Miller Levine Reading And Study Workbook A Answer Key
  - The Rise of Digital Reading Miller Levine Reading And Study Workbook A Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Miller Levine Reading And Study Workbook A Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Miller Levine Reading And Study Workbook A Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Miller Levine Reading And Study Workbook A Answer Key
  - Personalized Recommendations
  - Miller Levine Reading And Study Workbook A Answer Key User Reviews and Ratings
  - Miller Levine Reading And Study Workbook A Answer Key and Bestseller Lists

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  - Miller Levine Reading And Study Workbook A Answer Key Public Domain eBooks
  - Miller Levine Reading And Study Workbook A Answer Key eBook Subscription Services
  - Miller Levine Reading And Study Workbook A Answer Key Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - Miller Levine Reading And Study Workbook A Answer Key Compatibility with Devices
  - Miller Levine Reading And Study Workbook A Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Miller Levine Reading And Study Workbook A Answer Key
  - Highlighting and Note-Taking Miller Levine Reading And Study Workbook A Answer Key
  - Interactive Elements Miller Levine Reading And Study Workbook A Answer Key
8. Staying Engaged with Miller Levine Reading And Study Workbook A Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Miller Levine Reading And Study Workbook A Answer Key
9. Balancing eBooks and Physical Books Miller Levine Reading And Study Workbook A Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Miller Levine Reading And Study Workbook A Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Miller Levine Reading And Study Workbook A Answer Key
  - Setting Reading Goals Miller Levine Reading And Study Workbook A Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Miller Levine Reading And Study Workbook A Answer Key
  - Fact-Checking eBook Content of Miller Levine Reading And Study Workbook A Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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