

## CHAPTER SIX

- 6.1 a. AB: Heat liquid -  $-V \approx \text{constant}$   
BC: Evaporate liquid -  $-P'$  increases, system remains at point on vapor - liquid equilibrium curve as long as some liquid is present.  $T = 100^\circ\text{C}$ .  
CD: Heat vapor -  $T$  increases,  $V$  increases.

- b. Point B: Neglect the variation of the density of liquid water with temperature, so  $\rho = 1.00\text{ g/mL}$  and  $V_b = \underline{10\text{ mL}}$ .

Point C:  $\text{H}_2\text{O}$  (v,  $100^\circ\text{C}$ )

$$n = \frac{10\text{ mL}}{1\text{ mL}} \times \frac{1.00\text{ g}}{18.02\text{ g}} = 0.555\text{ mol}$$

$$P_C V_C = nRT_C \Rightarrow V_C = \frac{nRT_C}{P_C} = \frac{0.555\text{ mol}}{1\text{ atm}} \times \frac{0.08206\text{ L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \times \frac{373\text{ K}}{1} = \underline{17\text{ L}}$$

- 6.2 a.  $P_{\text{total}} = 243\text{ mm Hg}$ . Since liquid is still present, the pressure and temperature must lie on the vapor-liquid equilibrium curve, where by definition the pressure is the vapor pressure of the species at the system temperature.

- b. Assuming ideal gas behavior for the vapor,

$$m(\text{vapor}) = \frac{(3.000 - 0.010)\text{ L}}{(30 + 273.2)\text{ K}} \times \frac{\text{mol} \cdot \text{K}}{0.08206\text{ L} \cdot \text{atm}} \times \frac{243\text{ mm Hg}}{760\text{ mm Hg}} \times \frac{1\text{ atm}}{1} \times \frac{119.39\text{ g}}{\text{mol}} = 4.59\text{ g}$$

$$m(\text{liquid}) = \frac{10\text{ mL}}{1\text{ mL}} \times \frac{1.489\text{ g}}{1} = 14.89\text{ g}$$

$$m_{\text{total}} = m(\text{vapor}) + m(\text{liquid}) = \underline{19.5\text{ g}}$$

$$x_{\text{vapor}} = \frac{4.59}{19.48} = \underline{0.235\text{ g vapor / g total}}$$

- 6.3 a.  $\log_{10} p^* = 7.09808 - \frac{1238.71}{45 + 217} = 2.370 \Rightarrow p^* = 10^{2.370} = \underline{234.5\text{ mm Hg}}$

$$\text{b. } \ln p^* = -\frac{\Delta \hat{H}_v}{R} \frac{1}{T} + B \Rightarrow -\frac{\Delta \hat{H}_v}{R} = \frac{\ln(p_2^* / p_1^*)}{\frac{1}{T_2} - \frac{1}{T_1}} = \frac{\ln(760 / 118.3)}{\frac{1}{(77.0 - 273.2)\text{ K}} - \frac{1}{(29.5 - 273.2)\text{ K}}} = -4151\text{ K}$$

$$B = \ln(p_1^*) + \frac{\Delta \hat{H}_v / R}{T_1} = \ln(118.3) + \frac{4151\text{ K}}{(29.5 + 273.2)\text{ K}} = 18.49$$

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