

Experiment 3: Chemical Kinetics

Report Sheet/Post Lab Questions

Measure	Solution 1	Solution 2	Solution 3	† Temperature
Volume H ₂ O ₂ (mL)	5.0	10	5	5.0
Volume H ₂ O (mL)	15.0	10.0	5.0	15.0
Volume KI (mL)	10	10	20	10
Slope	0.1484	0.2677	0.3129	0.3018

1. Show your calculation for the slope.

$$m = 1$$

$$n = 2$$

2. Write the Rate Law for this reaction

$$\text{rate} = k[A][B]^2$$

3. What is the order of this reaction?

$$m+n = 3$$

4. Did the rate for the higher temperature increase or decrease?

higher

5. Look up the solubility of oxygen in water.

faster

6. From the solubility value, do you expect that its solubility can affect the experimental results? Explain why.

7. What are some possible errors in this experiment?

the time we get is no precise

8. Was iodide left in the reaction solution after it was allowed to go to completion? What can you conclude about iodide?

9. From the data collected could the activation energy for the reaction be calculated?

Kinetics Of A Reaction Lab Answers

**United States. Environmental
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