



Worksheet: Le Chatelier's Principle



1. Ammonia is produced commercially by the Haber reaction :



The formation of ammonia is favored by

- (A) an increase in pressure
(B) a decrease in pressure
(C) removal of $\text{N}_2 (\text{g})$
(D) removal of $\text{H}_2 (\text{g})$

Ans. _____

2. Given the closed system at equilibrium :



As the pressure on the system increases, the solubility of the $\text{CO}_2 (\text{g})$

- (A) decreases
(B) increases
(C) remains the same

Ans. _____

3. Consider the following equilibrium system :



Describe the effect that a decrease in volume would have on the position of equilibrium and the $[\text{H}_2]$ in the above system.

4. When the volume of the following mixture of gases is increased, what will be the effect on the equilibrium position?



5. Predict the effect of decreasing the container volume for each equilibrium.







Le Chateliers Principle Answer Key

Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name 'Mike Jess'.

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