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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	N-H	N-N
bond energy in kJ/mol	435	165

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an **N-O bond** in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (4 \times 435) + 1(165) \\
 & 4 \times 435 + 1(165) \quad \underline{2140} \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 & \text{bond broken} - \text{bond formed} = -90 \\
 & 2140 - x = -90 \\
 & x = 2140 + 90 \\
 & \quad \quad \quad \underline{2230} \text{ kJ}
 \end{aligned}$$
- Calculate the energy of **an N-O bond**.

$$\begin{aligned}
 & 6 \times (\text{N-O}) = 2230 \\
 & \text{N-O} = \frac{2230}{6} \quad \underline{372} \text{ kJ/mol}
 \end{aligned}$$

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