

You must show all your calculation set-ups to receive full credit. You must report all answers to the correct # of significant figures (sig. figs.).

1. Who used a cathode ray tube in his/her discovery of the electron?

- ☒ a. Robert Millikan
☐ b. Ernest Rutherford
☐ c. John Dalton
☒ d. J.J. Thomson

82
100

2. How many significant figures are in each of the following measurements?

- a. 0.0017 g 2 sig figs
 b. 1390 mL 3 sig figs
 c. 4.030 km 4 sig figs
 d. 200110 kg 5 sig figs

3. Which unit represents a microgram?

- ☒ a. mg
☒ b. μ g
☐ c. ng
☐ d. cg

4. The most unreactive group in the Periodic Table is/are the

- ☒ a. alkaline-earth metals
☐ b. metalloids
☐ c. halogens
☒ d. noble gases

5. The empirical formula of a compound of carbon and hydrogen is C_2H_3 . The molar mass is 28.06 g/mol. What is the molecular formula?

- ☒ a. C_4H_6
☐ b. C_3H_4
☐ c. C_2H_2
☐ d. C_2H_{10}

C: 12.01
 H: 1.01

$$28.06 = n(14.03)$$

$$n = 2$$

$$\frac{C \ 12.01}{H \ 2(1.01)}$$

$$14.03$$

6. Label each of the following elements as metal, nonmetal, or metalloid.

- a. chlorine nonmetal
 b. beryllium metal
 c. silicon metalloid
 d. boron metalloid

7. Write both the name and the symbol of the element in group 13 (3A) and period 3.

Aluminum (Al)

8. Calculate the percentage composition by mass (mass percent) of just the hydrogen in C_3H_8 .

$$\text{C}_3\text{H}_8 \rightarrow \text{total: } 44.11$$

$$C: 3(12.01) = 36.03 / 44.11 = 82\%$$

$$H: 8(1.01) = 8.08 / 44.11 = 18\%$$

Michigan Standardized College Chemistry Exam

Terry C. Jones



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Table of Contents Michigan Standardized College Chemistry Exam

1. Understanding the eBook Michigan Standardized College Chemistry Exam
 - The Rise of Digital Reading Michigan Standardized College Chemistry Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Michigan Standardized College Chemistry Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Michigan Standardized College Chemistry Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Michigan Standardized College Chemistry Exam
 - Personalized Recommendations
 - Michigan Standardized College Chemistry Exam User Reviews and Ratings
 - Michigan Standardized College Chemistry Exam and Bestseller Lists
5. Accessing Michigan Standardized College Chemistry Exam Free and Paid eBooks
 - Michigan Standardized College Chemistry Exam Public Domain eBooks
 - Michigan Standardized College Chemistry Exam eBook Subscription Services
 - Michigan Standardized College Chemistry Exam Budget-Friendly Options

6. Navigating Michigan Standardized College Chemistry Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - Michigan Standardized College Chemistry Exam Compatibility with Devices
 - Michigan Standardized College Chemistry Exam Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Michigan Standardized College Chemistry Exam
 - Highlighting and Note-Taking Michigan Standardized College Chemistry Exam
 - Interactive Elements Michigan Standardized College Chemistry Exam
8. Staying Engaged with Michigan Standardized College Chemistry Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Michigan Standardized College Chemistry Exam
9. Balancing eBooks and Physical Books Michigan Standardized College Chemistry Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Michigan Standardized College Chemistry Exam
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Michigan Standardized College Chemistry Exam
 - Setting Reading Goals Michigan Standardized College Chemistry Exam
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Michigan Standardized College Chemistry Exam
 - Fact-Checking eBook Content of Michigan Standardized College Chemistry Exam
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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