

- 2 In an experiment, powdered calcium hydroxide was added to dilute hydrochloric acid and the pH was measured.

The method used was

- step 1** measure  $200\text{ cm}^3$  dilute hydrochloric acid into a beaker  
**step 2** add  $0.1\text{ g}$  of powdered calcium hydroxide to the beaker  
**step 3** find the pH of the mixture  
**step 4** repeat steps 2 and 3 until the pH stops changing.

- (a) State what should be done after **step 2** to make sure that any reaction is complete.

(1)

- (b) Complete the word equation for the reaction.

(2)

calcium hydroxide + hydrochloric acid  $\rightarrow$  \_\_\_\_\_

- (c) Which row of the table shows the state symbols for powdered calcium hydroxide

|     |     | calcium hydroxide | hydrochloric acid |
|-----|-----|-------------------|-------------------|
| (1) | (2) | (3)               | (4)               |
| (1) | (2) | (3)               | (4)               |
| (1) | (2) | (3)               | (4)               |
| (1) | (2) | (3)               | (4)               |

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