

Chapter 5 Assessment

pages 166–169

Section 5.1

Mastering Concepts

- 34.** Define the following terms.
- frequency**
Frequency is the number of waves that pass a given point per second.
 - wavelength**
Wavelength is the shortest distance between equivalent points on a continuous wave.
 - quantum**
A quantum is the minimum amount of energy that can be lost or gained by an atom.
 - ground state**
An atom's ground state is its lowest allowable energy state.
- 35.** Arrange the following types of electromagnetic radiation in order of increasing wavelength.
- ultraviolet light
 - microwaves
 - radio waves
 - X rays
- d, X rays, a, ultraviolet light, b, microwaves, c, radio waves
- 36.** A gamma ray has a frequency of 2.88×10^{21} Hz. What does that mean?
- 2.88×10^{21} gamma-ray electromagnetic waves of this frequency pass a given point per second.
- 37.** What is the photoelectric effect?
- a phenomenon in which a metal emits electrons when light of a sufficient frequency shines on it
- 38. Neon Sign** How does light emitted from a neon sign differ from sunlight?
- The light from a neon sign contains only certain visible colors, while sunlight contains the full spectrum of colors.
- 39.** Explain Planck's quantum concept as it relates to energy lost or gained by matter.
- According to Planck, for a given frequency, ν , matter can emit or absorb energy only in discrete quanta that are whole-number multiples of $h\nu$, where h is Planck's constant.
- 40.** How did Einstein explain the photoelectric effect?
- He proposed that photons must have a certain minimum energy level, or threshold, value to cause the ejection of a photoelectron.
- 41. Rainbow** What are two differences between the red and green electromagnetic waves in a rainbow?
- The red waves have a longer wavelength and a lower frequency.
- 42. Temperature** What happens to the light emitted by a heated, glowing object as its temperature increases?
- The color of the light changes as the object acquires more and more energy.
- 43.** What are three deficiencies of the wave model of light related to light's interaction with matter?
- The wave model does not explain the photoelectric effect, atomic emission spectra, and why matter emits different frequencies of light at different temperatures.
- 44.** How are radio waves and ultraviolet waves similar? How are they different?
- Both types of waves travel at the same speed in a vacuum, 3.00×10^8 m/s. Radio waves have a longer wavelength and lower frequency than ultraviolet waves.

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John G. Truxal



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