

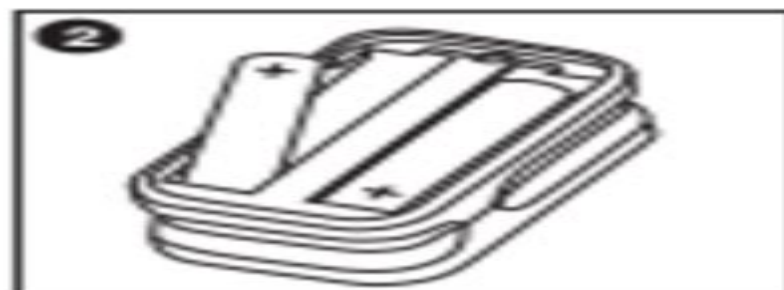
# BRAUN

Pulse oximeter 1  
Pulsoximeter 1

YK-81C

Installing batteries  
Batterien einlegen

## Initial set up Erste Einrichtung



Pulse oximeter 1 – Instalning af batterier – Indledende opsætning  
Pulsoxímetro 1 – Instalación de las baterías – Configuración inicial  
Pulsoxímetri 1 – Paristojen asentaminen – Alkasetus  
Oxymètre de pouls 1 – Installation des piles – Configuration initiale  
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Oxímetro de pulso 1 – Colocar as pilhas – Instalação inicial

موصل التأكسج النبضي 1 - تركيب البطاريات - الإعداد الأولي

Pulse oximeter 1 – Sättas i batterier – Initial installation

Pulse oximeter cihazı 1 – Pillerin takılması – İlk kurulum

# Manuale Uso Braun L918

**Lingjun Ying**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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web 1 practice problem balance the redox reaction in basic solution  $\text{NO}_2^- + \text{Al} + \text{S} + \text{NH}_3 \rightarrow \text{AlO}_2^- + \text{NO}_3^- + \text{S} + \text{NH}_4^+$  practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of  $\text{OH}^-$  br 2 l bro 3 aq br aq previous topic next topic

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web google classroom for the following redox equation  $\text{MnO}_4^- + \text{SO}_2 + 2\text{H}^+ \rightarrow \text{Mn}^{2+} + \text{H}_2\text{SO}_4$  what will be the stoichiometric coefficient of  $\text{H}_2\text{SO}_4$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN} + 2\text{e}^-$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- \rightarrow 2\text{AuCN} + 2\text{e}^-$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{OH}^-$  3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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reaction in acid solution using the half reaction method 8 h<sub>2</sub>o<sub>2</sub> cr<sub>2</sub>o<sub>7</sub> 2 9 teo<sub>3</sub> 2 n<sub>2</sub>o<sub>4</sub> o<sub>2</sub> cr<sub>3</sub> te 10 reo<sub>4</sub> 11 pbo<sub>2</sub> io i<sub>2</sub> pb<sub>2</sub> 12 as no<sub>3</sub> io<sub>3</sub> re io<sub>3</sub> h<sub>2</sub>aso<sub>4</sub>

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h<sup>+</sup> 2no<sup>2-</sup> 2 c 6 h 6 o 6 2h<sup>+</sup> 2 o<sup>2-</sup> 2no choose 1 answer c 6 h 8 o 6 and no<sup>2-</sup> a c 6 h 8 o 6 and no<sup>2-</sup> no 2 and c 6 h 8 o 6 b no<sup>2-</sup> and c 6 h 8 o 6 h<sup>+</sup> and no<sup>2-</sup> c h<sup>+</sup> and

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