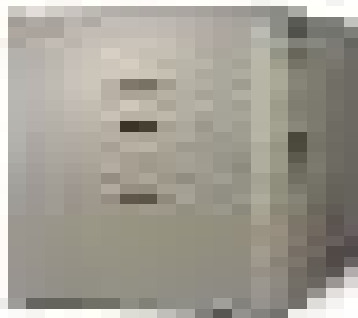


# LABORATORY

EXPERIMENT 1  
PURPOSE  
OBJECTIVES  
THEORY  
PROCEDURE



RESULTS  
DISCUSSION  
CONCLUSION  
REFERENCES

INTRODUCTION  
The purpose of this experiment is to determine the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction is exothermic and produces iodine and water. The rate of reaction is measured by the time taken for a fixed volume of iodine to be produced. The experiment is carried out at different temperatures and the results are plotted on a graph of time against temperature. The graph shows that the rate of reaction increases with temperature. This is due to the fact that at higher temperatures, the molecules have more kinetic energy and are therefore more likely to collide and react. The experiment also demonstrates the effect of a catalyst on the rate of reaction. Potassium iodide acts as a catalyst for the reaction and increases the rate of reaction. This is because it provides an alternative reaction pathway with a lower activation energy. The experiment is carried out in a series of test tubes and the results are recorded in a table. The table shows that the rate of reaction increases with temperature and that the use of a catalyst also increases the rate of reaction. The experiment is a good example of how temperature and catalysts can affect the rate of a chemical reaction.

# Luxaire Furnace Manual

**EW Minium**



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