CHEMISTRY - (Theory)

Dec. 2022 - 2 hours



Name	Index Number
Candidate's Signature	Date

Instructions to candidates

- (a) Write your name and index number in the spaces provided above.
- (b) Sign and write the date of examination in the spaces provided above.
- (c) Answer all the questions in the spaces provided in the question paper.
- (d) Non-programmable silent electronic calculators and KNEC mathematical tables may be used.
- (e) All working must be clearly shown where necessary.
- (f) This paper consists of 16 printed pages.
- (g) Candidates should check the question paper to ascertain that all the pages are printed as indicated and that no questions are missing.
- (h) Candidates should answer the questions in English.

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Question	Maximum Score	Candidate's Score
1	11	reach the second
2	/ 13	15.30
3	12	100
4	13	1000
5	11	
Co. Carrows	10	25.00
7	10	7
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Pat Carlson Willis

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model Bohr's atomic model defects boiling point and external pressure boiling points bond formation Boyle's law charge to mass ratio of electron Charles s law chemical bonding chemical combinations chromatography classification of solids combustion analysis covalent radius covalent solids crystal lattice crystallization crystals and classification cubic close packing diamond structure diffusion and effusion dipole forces dipole induced dipole forces discovery of electron discovery of neutron discovery of proton dual nature of matter dynamic equilibrium electron affinity electron charge electron distribution electron radius and energy derivation electron velocity electronic configuration of elements empirical formula energy changes and intermolecular attractions energy of revolving electron experimental techniques filter paper filtration crucibles fundamental particles gas laws gas properties graham s law grahams law of diffusion Heisenberg s uncertainty principle hexagonal close packing higher ionization energies hydrogen bonding hydrogen spectrum ideal gas constant ideal gas density ideality deviations intermolecular forces ionic radius ionization energies ionization energy isotopes kinetic interpretation of temperature kinetic molecular theory of gases Lewis concept liquefaction of gases liquid crystals liquids properties London dispersion forces magnetic quantum number mass of electron mass spectrometer metallic crystals properties metallic solids metals structure molar volume molecular ions molecular solids molecules moles Moseley law neutron properties non ideal behavior of gases orbital concept partial pressure calculations phase changes energies photons wave number Planck's quantum theory plasma state positive and negative ions pressure units properties of cathode rays covalent crystals properties of crystalline solids properties of positive rays quantum numbers quantum theory relative abundance Rutherford model of atom shapes of orbitals solid iodine structure solids properties solvent extraction spectrometer spin quantum number states of matter stoichiometry sublimation thermometry scales types of solids unit cell van der Waals equation vapor pressure and spectrum **Chemistry Exam Answers** Sarah Knowlton, Higher School Certificate Chemistry Exam Questions with Answers and Worked Solutions Vera Liondas, Graham Ware, 1982-01 How to Pass KCSE Chemistry Practical Pat Carlson Willis, 2024-01-10 Embark on a captivating journey through the realm of chemistry with How to Pass KCSE Chemistry Practical Ultimate Companion for Success Born out of over 35 years of passionate teaching experience this comprehensive guide is meticulously designed to demystify and enhance the learning experience for students preparing for the Kenyan Certificate of Secondary Education KCSE Chemistry Practical exam More than just a textbook it is a friendly mentor offering a wealth of insights practical tips and engaging content to ensure success in mastering the nuances of practical chemistry The book covers an array of captivating topics starting with qualitative analysis Dive into the fascinating world of identifying ions and compounds through experimental techniques with step by step instructions that transform complex procedures into accessible and enjoyable learning experiences From understanding the intricacies of titration to mastering the art of precise measurements the book provides a comprehensive guide to ensure students approach this essential technique with confidence and skill Rates of reaction take center stage as the book explores

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