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**0620/31**

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# May June 2013 Chemistry Paper 61 0620

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$$\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$$

$$2\text{C}_6\text{H}_{10}\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + \text{GaBr} + \text{NaBr} + \text{SnO}$$

$$4\text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + \text{SO}_3 \rightarrow 3\text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$$
 using the equation from problem 2 above answer the following

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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$3\text{HBr} + 1\text{KHCO}_3 + 1\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2 + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{SO}_3 + 6\text{NaBr} + 3\text{SnO} + 2\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$$
 using the following equation  

$$2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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$$\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn} + \text{CuCl}_x$$
 2 ZnCl<sub>x</sub> 2 Cu how many moles of Ce ZnCl<sub>2</sub> ZnCl<sub>x</sub> 2 will be produced from 23.0 g of Ce Zn Zn assuming Ce CuCl<sub>2</sub> CuCl<sub>x</sub> 2 is available in excess

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web jul 22 2013 balance the following equations  

$$\text{Br}_2 + \text{N}_2 + 3\text{F}_2 + 2\text{NF}_3 + 2\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 + 12\text{CO}_2 + 10\text{H}_2\text{O} + 3\text{HBr} + 1\text{KHCO}_3 + 1\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2 + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{SO}_3 + 6\text{NaBr} + 3\text{SnO} + 2\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$$
 using the following equation  

$$\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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