

Nace Chapters 1-5 Exam 3 Study Guide With 100% Complete Solution

Corrosion description

Is usually described by its results. Familiar terms include rust, scaling, discoloration, oxidation, pitting, etc.

Corrosion

The deterioration of a substance, usually a metal, or its properties because of a reaction with its environment.

The most common product of corrosion is an

Oxide of iron "rust"

Passivation

The formation of a thin, tightly adherent surface layer. Formed from the corrosion product, that slows the rate of corrosion.

All corrosion of iron at normal ambient conditions is
an electrochemical process.

How do electrons and ions travel,

Electrons travel through metal, ions travel through an electrolyte.

Corrosion is established as

DC circuits.

The greater the flow in the corrosion circuit,

The greater the metal loss.

Anode

is the part of the metal that corrodes (i.e., dissolves in the electrolyte).

The metal that dissolves does so in the form of

positively charged ions.

Metal deterioration occurs at the

Anode. It is the portion of the cell that where metallic iron is first converted to another substance.

Cations

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